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Got questions or suggestions? - Email us at - Terms of Service - Privacy Policy: out how your app performs in app stores Find a Tutor Revision Resources PMT Resources Predicted Papers Revision Resources Revision Courses Teacher Resources Blog PMT Blog Students Parents Teachers Tutors STEM This website contains the notes I wrote when studying for A Level Chemistry, Physics, Mathematics and Further Maths, organised by module/exam. The physics and chemistry ones should cover the entire syllabus and also include about 2100 flashcards. (A Levels are British high school exams: usually students take 3 or 4, age 16-18). My GCSE Music, Biology, Chemistry, and Physics notes, written when I was a GCSE student in 2015. GCSEs have since changed a lot, but you may find this useful. Hi, my name is James (I'm sure you've already figured that out...). I co-own a web design company. This is one of my first websites from when I was studying for GCSEs and A Levels. Feel free to email me at or SMS/WhatsApp +44 7377 565962 This website is designed to work on nearly all public library and school computers, phones, tablets, and printed on paper. Please contact me if you have any technical issues.Note: I am not a teacher or scientist, only an ex-student. I wrote the content myself in preparation for the exams, I chose to share it for free online since many students find it helpful. Encouraging and passionate Biology and Chemistry tutor with 6 years experience. MSc in Chemistry from Imperial College. For each of the papers below, there are revision notes, summary sheets, questions from past exam papers separated by topic and other worksheets. Page 2Isotopes are atoms of the same element that contain the same number of protons and electrons but a different number of neutronsThe name of an isotope is the chemical symbol (or word) followed by a dash and then the mass numberE.g. carbon-12 and carbon-14 are isotopes of carbon containing 6 and 8 neutrons respectivelyThe chemical symbol of an isotope will still be shown in the usual form, as shown:Isotopes of the same element have the same chemical symbol with a different mass numberFor chemical symbols, isotopes have:The same chemical symbole.g. isotopes of chlorine all have the symbol ClThe same atomic numbere.g. isotopes of chlorine all have an atomic number of 17A different mass / nucleon numbere.g. the chlorine-35 isotope has a mass number of 35, while the chlorine-37 isotope has a mass number of 37All three hydrogen isotopes have the chemical symbol H, the same number of protons but a different number of neutronsIsotopes have similar chemical properties but different physical propertiesThe chemical behaviour of an atom depends on its electron configuration, especially the electrons in the outer (valence) shellThis is because all isotopes of the same element have the same number of electrons and the same electron arrangement, they:Form the same types of bondsReact with the same elementsShow identical chemical reactivityAlthough isotopes have identical chemical properties, they differ in physical properties because they have different numbers of neutronsNeutrons add mass to the nucleus but carry no charge and don't influence chemical bondingAs a result, isotopes have:Different relative atomic massesSlight differences in densityThese small differences in physical properties become important in processes like:Mass spectrometry (detecting different isotopes by mass)Isotope separation (e.g. in nuclear chemistry)Medical imaging (radioisotopes rely on differences in physical properties)Did this page help you?

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